Structure of \((\text{Ta}_2\text{O}_5)_x(\text{SiO}_2)_{1-x}\) xerogels \((x = 0.05, 0.11, 0.18, 0.25\) and 1.0\) from FTIR, \(^{29}\text{Si}\) and \(^{17}\text{O}\) MAS NMR and EXAFS

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A combination of \(^{29}\text{Si}\) and \(^{17}\text{O}\) MAS NMR, EXAFS and FT-IR spectroscopy has been used to study the atomic structure of \((\text{Ta}_2\text{O}_5)_x(\text{SiO}_2)_{1-x}\) \((x = 0.05, 0.11, 0.18\) and 0.25\) xerogels prepared by reacting partially-hydrolysed tetraethyl orthosilicate with tantalum(v) ethoxide. Amorphous tantalum-\(\alpha\)-Ta\(_2\)O\(_5\) xerogels have also been prepared and their structures studied in detail for the first time. Results have shown that in all these materials, Ta adopts predominantly 5-fold coordination with respect to oxygen. For the mixed oxide xerogels, partial phase separation of the two component oxides occurs for \(x > 0.11\).

**Introduction**

\((\text{Ta}_2\text{O}_5)_x(\text{SiO}_2)_{1-x}\) mixed oxide materials have attracted interest because of their technologically useful properties. The sol–gel route from alkoxide precursors\(^+\) has been used to prepare such materials because it offers a low-temperature synthesis to materials with high purity and homogeneity. \((\text{Ta}_2\text{O}_5)_x(\text{SiO}_2)_{1-x}\) mixed oxides prepared this way possess high refractive indices rendering them suitable for optical components such as micro-lenses and optical fibres.\(^6\) The high refractive index of a given sample has been shown to depend upon the level of atomic mixing of the two oxides, a property which in turn is sensitive to the exact method of synthesis. These materials also exhibit catalytic activity which derives from the acidity of the Ta–O–Si linkages present in their structures.\(^3\) Again, preparation and composition are important in producing good catalysts with high surface areas and a high level of atomic mixing. In general, the useful properties of sol–gel derived \((\text{Ta}_2\text{O}_5)_x(\text{SiO}_2)_{1-x}\) mixed oxides are reduced by any phase separation of the two component oxides. Clearly, a sensitive approach to optimizing the properties of these materials would be to study the atomic level structure in order to understand how the oxides mix from a structural point of view and to gain information on the mechanism of phase separation.

In this paper, the atomic structure of a series \((\text{Ta}_2\text{O}_5)_x(\text{SiO}_2)_{1-x}\) xerogels of various compositions and thermal histories have been studied using a comprehensive advanced methodology involving a combination of probe techniques.

**Experimental**

**Sample preparation**

The samples were prepared using the sol–gel route from the following precursors: tetraethyl orthosilicate, TEOS (Aldrich, 98%), and tantalum(v) ethoxide, Ta(OEt)\(_5\) (Aldrich, 99.98%). HC\(_2\)F (Fisons) was used as a catalyst to promote the hydrolysis and condensation reactions and ethanol (Aldrich, 99.9%) was used as a mutual solvent.

The method of Yoldas\(^4\) was used to promote homogeneity within the mixed oxide samples. This involved prehydrolysis of the TEOS to maximise the number of Si–OH groups before mixing with the more reactive Ta(OEt)\(_5\) precursor; the aim being to encourage Ta–O–Si bonding as opposed to Ta–O–Ta bonding which could lead to phase separation. The chosen prehydrolysis conditions were TEOS : EtOH : H\(_2\)O in a 1 : 1 : 1 molar ratio in the presence of HCl (pH = 1), stirring for 2 hours. The appropriate quantity of Ta(OEt)\(_5\) was then slowly added to the prehydrolysed TEOS solution while stirring. After allowing several minutes for the oxide precursors to mix, water was added such that the overall water alkoxide molar ratio was 2. The resulting clear sol was then left to gel; this typically took a few days, depending on composition.

An amorphous tantalum oxide, \(\alpha\)-Ta\(_2\)O\(_5\), sample was also prepared so that the structural environment of Ta could be compared with that in the mixed oxides. This was prepared by slowly adding an acidified H\(_2\)O/EtOH (1 : 8 molar ratio) solution to an 0.5 M solution of Ta(OEt)\(_5\) in EtOH while stirring.\(^5\) A milk-white gel formed almost immediately.

It should be noted that all reagents were loaded in a dry box under a N\(_2\) atmosphere and transferred using syringes to avoid absorption of moisture from the atmosphere. Samples for \(^{17}\text{O}\) MAS NMR were prepared using 10 mol\% \(^{17}\text{O}\) enriched water (D-CHEM).

All samples were aged for one week before being air dried for several days, finely ground and then pumped under vacuum for 24 hours to remove any excess solvent. Heat treatments were performed at a heating rate of 5 °C min\(^{-1}\) with each temperature maintained for 2 hours. Samples were prepared with nominal compositions of \((\text{Ta}_2\text{O}_5)_x(\text{SiO}_2)_{1-x}\) where \(x = 0.05, 0.11, 0.18, 0.25\) and each composition heated to 250, 500 and 750 °C. The \(\alpha\)-Ta\(_2\)O\(_5\) sample was calcined at 250 and 750 °C. \(^{17}\text{O}\) enriched samples were prepared for the compositions \(x = 0.05, 0.25\) and 1. These enriched samples were heat treated at 125, 250, 350, 500 and 750 °C under a N\(_2\) atmosphere in order to prevent the loss of \(^{17}\text{O}\).

**FT-IR spectroscopy**

Infrared spectra were recorded in diffuse reflectance mode on a Biorad FTIR 175C spectrometer controlled by Win-IR software. Samples were diluted (1 : 10 by weight) in dry KBr and scanned over the range 4000–400 cm\(^{-1}\) with a resolution of 4 cm\(^{-1}\). Each spectrum was the result of summing 64 scans.

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MAS NMR

The MAS NMR spectra were acquired on a Chemagnetics CMX300 Infinity spectrometer. The $^{29}$Si spectra were collected at 59.6 MHz using MAS at typically 6 kHz with a 2.0 $\mu$s (~30° tip angle) pulse and a 20 $s$ recycle delay. A 7.5 mm Chemagnetics double-bearing probe was used. The pulse delay was sufficient to prevent saturation. The spectra were referenced externally to Zeolite A (~89.7 ppm). Each spectrum was the result of summing ~3000 scans. The $^{17}$O NMR spectra from the $(Ta_2O_5)_x(SiO_2)_{1-x}$ xerogels were collected at 40.7 MHz under MAS at typically 15 kHz with a recycle delay of 1–2.5 $s$ using a Chemagnetics 4 mm double-bearing probe. The delay was sufficient to prevent saturation. A 90°–c–180° echo sequence was applied with a short $\tau$ delay of 71.4 $\mu$s to overcome problems of probe ringing, thereby allowing the spectra to be phased correctly.

For the pure sol–gel prepared $Ta_2O_5$, the $^{17}$O MAS NMR data were collected at a resonance frequency of 32.1 MHz using a Bruker 4 mm double-bearing MAS probe applying the same experimental approach and very similar conditions to those used for the mixed oxide gels.

All the $^{17}$O MAS NMR spectra were referenced externally to $H_2O$ (0 ppm). Typically each spectrum was the result of 50 000–100 000 co-added scans and took approximately 24 $h$ to collect.

X-Ray diffraction

The powder X-ray diffraction pattern of the pure $Ta_2O_5$ xerogel heated to 750 °C was collected on a Philips PW1050 diffractometer using Cu-K$_{\alpha}$ radiation (\(\lambda = 1.541\) Å). A range of 29 = 20–70° was scanned at a rate of 1° min$^{-1}$ with a step size of 0.02°. An estimate of particle size (S) was made using the Scherrer equation, $S = K\lambda /\beta \cos \theta$, where $K$ is a constant (0.89) and $\beta$ is the full width at half height of a diffraction peak at angle $\theta$.

EXAFS spectroscopy

Ta L$_{3,4}$-edge EXAFS data were collected in transmission mode at room temperature on station 7.1 at the Daresbury Laboratory SRS using a Si [111] crystal monochromator and 50% harmonic rejection. Ioniisation chambers, filled with a mixture of Ar/He or Kr/He at appropriate partial pressures to optimise detector sensitivities, were placed in the beam path before and after the sample. Finely-ground samples were diluted in polyethylene (Aldrich, spectrophotometric grade) and pressed into pellets to give a satisfactory edge jump and absorption. Typically each EXAFS spectrum was the sum of 45 mm scans.

Results and data analysis

FTIR spectroscopy

Fig. 1 (A) and (B) show the IR spectra of the $(Ta_2O_5)_x(SiO_2)_{1-x}$ xerogels unheated and after heating to 750 °C, respectively. Also shown, for comparison, is the spectrum from an $a$-$Ta_2O_5$ xerogel after heating to 250 °C. The peaks in these spectra have been assigned according to the literature.4 $-$ 12 The band at 3590 cm$^{-1}$ is mainly due to stretching vibrations involving Si–OH groups and the band at 3350 cm$^{-1}$ corresponds to adsorbed water. The band at 1610 cm$^{-1}$ is assigned to the H–O–H bending vibration of water. The peaks at 1160, 1070 and 900 cm$^{-1}$ are characteristic of a silica network: the bands at 1160 cm$^{-1}$ and 1070 cm$^{-1}$ are ascribed to the TO and TO components of the asymmetric stretch of the SiO$_4$ unit, respectively, and the feature at 800 cm$^{-1}$ is due to the symmetric stretch of the SiO$_4$ unit. The band at about 960 cm$^{-1}$ is best assigned to the stretching vibration of Si–O–Si groups$^{12}$ adjacent to different countercations; in this case, this band has contributions from both Si–OH and Si–O–Ta vibrations.

MAS NMR

The $^{29}$Si MAS NMR spectra were deconvoluted by Gaussian fitting using SpinSight 3.02 software; three signals could be distinguished at ~93, ~101 and ~109 ppm with varying intensities. Each resonance represents a specific degree of Si–O–Si polymerisation. The ~93 ppm signal comes from Si sites in a Q$^4$ configuration, and the ~101 and ~109 ppm signals come from Q$^2$ and Q$^3$ configurations, respectively (Q$^4$ stands for a SiO$_4$ unit with $a$ bridging oxygen). The results of the Gaussian fitting are summarised in Table 1. Fig. 2 shows the $^{29}$Si NMR data and peak deconvolution for the $(Ta_2O_5)_{0.95}(SiO_2)_{0.05}$ sample and clearly illustrates a shift to higher average $Q$ species with heat treatment.

The $^{17}$O MAS NMR spectra from pure $Ta_2O_5$ xerogels are shown in Fig. 3, and the major resonances listed in Table 2. In the dried, unheated xerogel two relatively narrow resonances at ~270 and ~430 ppm are observed along with some much broader underlying intensity that extends to ~0 ppm. On heating, the spectrum largely remains unchanged after heating up to 350 °C apart from a small positive shift of the peak that starts off at ~270 ppm (Table 2) and a change in the relative intensity. After heating to 500 °C there is a more significant change of the structure with some broader underlying intensity, and then at 750 °C the peak at 305 ppm becomes dominant.

This resonance is due to Si–O–Si bridged$^{11}$ with a contribution
from Si–OH groups. This signal has spinning sidebands from the MAS process at about 360 and −360 ppm. The resonances at −285 and −445 ppm observed for the (TaO)x0.25(SiO)0.75 samples are at similar chemical shifts to those in the pure Ta2O5 xerogel and hence can be assigned to Ta–O–Ta bonding. By comparison to the resonances due to Si–O–Si and Ta–O–Ta oxygen environments, the signal at around 160 ppm observed for both compositions has been assigned to Ta–O–Si bridges. Both the Ta–O–Ta and Ta–O–Si resonances display a small change to higher chemical shift with increasing temperature of the heat treatment.

### X-Ray diffraction

The powder X-ray diffraction pattern from the pure Ta2O5 xerogel heated to 750 °C indicates that crystallization has occurred by this calcination temperature. The powder pattern obtained matches the JCPDS card number 19-1299 for Ta2O5.

An average particle size of ~400 Å was determined from the widths of the Bragg peaks using the Scherrer equation.

### EXAFS spectroscopy

A heuristic version of the equation for the interpretation of EXAFS data is

\[
\chi(k) = A F A C \sum_{j} \frac{N_j}{k R_j^2} f(r, k, R_j) e^{-2k R_j} e^{-2\sigma^2 k^2} \sin[2k R_j + 2\delta(k) + \psi(k, R_j)]
\]

where \( \chi(k) \) is the magnitude of the X-ray absorption fine structure as a function of the photoelectron wave vector \( k \). \( F A C \) is the proportion of electrons that perform an EXAFS-type scatter. \( N_j \) is the coordination number and \( R_j \) is the interatomic distance for the \( j \)th shell. \( \delta(k) \) and \( \psi(k, R_j) \) are the phase shifts experienced by the photoelectron, \( f(r, k, R_j) \) is the amplitude of the photoelectron backscattering and \( \lambda(k) \) is the electron mean free path; these are calculated within EXCURV98. The Debye–Waller factor is \( A = 2\sigma^2 \) in EXCURV98.

The programs EXCALIB, EXBACK and EXCURV98 were used to extract the EXAFS signal and analyse the data. Least squares refinements of the structural parameters of our samples were carried out against the \( k^2 \) weighted EXAFS signal.
Table 2  $^{17}$O MAS NMR chemical shifts

<table>
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<th>Sample</th>
<th>Heat treatment/$^\circ$C</th>
<th>Peak positions/ppm (± 5 ppm)</th>
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<tr>
<td>$\text{TasO}_x$ xerogel</td>
<td>none</td>
<td>428</td>
</tr>
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<td></td>
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</tr>
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<td></td>
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<tr>
<td></td>
<td>750</td>
<td>455</td>
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<tr>
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</tr>
<tr>
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</tr>
<tr>
<td></td>
<td>750</td>
<td>447</td>
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<tr>
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</tr>
<tr>
<td></td>
<td>125</td>
<td>431</td>
</tr>
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</tr>
<tr>
<td></td>
<td>750</td>
<td>447</td>
</tr>
</tbody>
</table>

To minimise the fit index, FI,

$$FI = \sum \left| k^3 (x^T_k - x^R_k) \right|^2$$

where $x^T_k$ and $x^R_k$ are the theoretical and experimental EXAFS, respectively. The results of the refinements are reported in terms of the discrepancy index, $R_d$,

$$R_d (\%) = \frac{\int \left| \chi^2(k) - \chi^R(k) \right| k^3 \, dk}{\int \chi^R(k) k^3 \, dk}$$

Crystalline $\text{TasO}_x$, or tannite (Strem), was run as a reference material to check the validity of our data analysis and also to allow refinement of the parameter AFAC. The results obtained from this standard are summarised in Table 3. The parameters were obtained by fitting the EXAFS data over the range 2.5–16.0 Å$^{-1}$ (i.e. equivalent to 5–32 Å$^{-1}$ in conventional diffraction formalism). In these refinements, the coordination numbers were fixed at values consistent with the crystallographic data available. It should be noted that the parameters derived from the crystallographic data included in Table 3 are much simplified: the crystal structure of tannite contains Ta atoms in thirteen inequivalent sites. The parameters shown are a weighted average over the 24 Ta atoms within the unit cell. Since the 13 Ta environments contain such a large range of Ta–O and Ta–Ta distances, only the three most significant shells have been used to model the EXAFS data: any other shells with $R > 4$ Å have such large Debye–Waller factors that they would not contribute significantly to the EXAFS.

![Fig. 4](image1.png) $^{17}$O MAS NMR spectra of a $(\text{TasO})_{0.25}(\text{SiO}_2)_{0.75}$ xerogel after various heat treatments: a) no heat treatment, b) 125 $^\circ$C, c) 250 $^\circ$C, d) 350 $^\circ$C, e) 500 $^\circ$C and f) 750 $^\circ$C (peaks at 360 and $-360$ ppm are spinning side bands).

![Fig. 5](image2.png) $^{17}$O MAS NMR spectra of a $(\text{TasO})_{0.80}(\text{SiO}_2)_{0.20}$ xerogel after various heat treatments: a) no heat treatment, b) 125 $^\circ$C, c) 250 $^\circ$C, d) 350 $^\circ$C, e) 500 $^\circ$C and f) 750 $^\circ$C (peaks at 360 and $-360$ ppm are spinning side bands).
signal. The errors are the statistical uncertainties calculated by EXCURV98; the real uncertainties will be higher. A value of 0.74 for AFAC was derived from the data from tantalite; this was used throughout analysis of the data from our xerogel samples.

Also included in Table 3 are the structural parameters derived from the analysis of the Ta L\textsubscript{III} edge EXAFS data from the pure Ta\textsubscript{2}O\textsubscript{5} xerogel both before and after heat treatment to 250 and 750 °C. The data for the unheated and 250 °C samples have been fitted using a single shell, i.e. one Ta–O correlation; attempts to fit a Ta–Ta shell resulted in an unrealistic distance. A two shell fit was used for the data from the 750 °C sample which exhibited a strong peak at ~3.9 Å in the Fourier transform. Fig. 6 shows the EXAFS data and their Fourier transform together with the calculated fit for the unheated a-Ta\textsubscript{2}O\textsubscript{5} xerogel.

The EXAFS derived structural parameters for the (Ta\textsubscript{2}O\textsubscript{5})\textsubscript{2}–(SiO\textsubscript{2})\textsubscript{2}– xerogel samples are shown in Table 4. These data were best fitted with a model consisting of Ta–O and Ta–Si shells. The errors in Table 4 are the statistical uncertainties calculated by EXCURV98. Fig. 7 and 8 show the EXAFS data and their Fourier transforms together with the calculated fits for the two extremes of composition, (Ta\textsubscript{2}O\textsubscript{5})\textsubscript{0.99}(SiO\textsubscript{2})\textsubscript{0.95} and (Ta\textsubscript{2}O\textsubscript{5})\textsubscript{0.25}(SiO\textsubscript{2})\textsubscript{0.75}, after heat treatment to 750 °C.

**Discussion**

As can be seen from Table 1, the \textsuperscript{29}Si MAS NMR results from the (Ta\textsubscript{2}O\textsubscript{5})\textsubscript{3}–(SiO\textsubscript{2})\textsubscript{2}– xerogels exhibit a general trend towards more polymerised Q\textsuperscript{2} species as the temperature of the heat treatment increases. This result is expected, since all the samples are composed predominantly of SiO\textsubscript{4} and it is well known that when a silica sol-gel is heated to progressively higher temperatures, further condensation and the resultant increase in cross-linking leads to an increase in the proportions of higher Q species. The point to note with these NMR results is that there is a distribution of silica species. Table 1 clearly shows that as the tantalum dopant level is increased, the extent of condensation of the silica network is reduced. This effect is very similar to that of Zr in (Zr\textsubscript{2}O\textsubscript{5})\textsubscript{3}–(SiO\textsubscript{2})\textsubscript{2}– xerogels.\textsuperscript{29} In such materials, the presence of silicon in Si(SiO\textsubscript{2})\textsubscript{2}ZrO, \textsubscript{2}Q\textsubscript{10′} and Si(SiO\textsubscript{2})\textsubscript{2}ZrO\textsubscript{2} \textsubscript{2}Q\textsubscript{20′} environments contribute to the signals in the \textsuperscript{29}Si NMR spectrum at ~101 and ~93 ppm, respectively. Zirconium clearly acts as a network modifier. This effect is due to the high field strength\textsuperscript{23} of Zr\textsuperscript{4+} ions which causes them to adopt higher coordination numbers with respect to oxygen. Hence, Zr\textsuperscript{4+} cannot substitute tetrahedrally into the silica network. The behaviour of the Ta\textsuperscript{5+} in the samples studied here is also that of a network modifier with the NMR signals at ~83, ~93 and ~101 ppm having contributions from Si(SiO\textsubscript{2})\textsubscript{2}TaO, \textsubscript{2}Q\textsubscript{10′}, Si(SiO\textsubscript{2})\textsubscript{2}TaO\textsubscript{2} \textsubscript{2}Q\textsubscript{20′} and Si(SiO\textsubscript{2})\textsubscript{2}TaO\textsubscript{2} \textsubscript{2}Q\textsubscript{20′}. The overall trend in the data of a shift of the Q species distribution to more Q\textsuperscript{2} and Q\textsuperscript{3} species with higher Ta concentration is a direct consequence of the network disrupting effect of Ta. This effect is emphasized by the widths of the NMR peaks which show a general increase with increasing Ta content.

The FTIR results shown in Fig. 1(A) and (B) give an insight into the interaction between the two oxides present in the (Ta\textsubscript{2}O\textsubscript{5})\textsubscript{2}–(SiO\textsubscript{2})\textsubscript{2}– xerogels. In both sets of spectra, the strength of the silica network vibrations at 1160, 1070, and 800 cm\textsuperscript{−1} are reduced as the concentration of Ta is increased. This effect is most dramatic for the symmetric stretch of SiO\textsubscript{2} which is only just detectable in the (Ta\textsubscript{2}O\textsubscript{5})\textsubscript{0.25}(SiO\textsubscript{2})\textsubscript{0.75} sample both before heat treatment and after heating to 750 °C. The reduction in the strength of the silica network vibrations is due to the network modifying properties of Ta. These infrared results therefore support the conclusions drawn from the \textsuperscript{29}Si NMR data.

The infrared results also yield qualitative information on the extent of Ta–O–Si bonding within the structures of these xerogels. This information can be derived by examining the absorption band at 960 cm\textsuperscript{−1} which is ascribed to a combination of Si–OH and Ta–O–Si vibrations. Considering the IR spectra from the unheated xerogels, with the exception of the (Ta\textsubscript{2}O\textsubscript{5})\textsubscript{0.25}(SiO\textsubscript{2})\textsubscript{0.75} sample, all of the samples contain similar amounts of hydroxyl groups, as seen from the intensities of the
bands at 3590 cm⁻¹. Hence, comparison of the intensities at 960 cm⁻¹ should give an indication of the relative amount of Ta-O-Si bonding in each sample. Clearly, for the x = 0.05, 0.11 and 0.18 samples, as the concentration of Ta increases so does the intensity of the 960 cm⁻¹ band relative to the intensity of the SiO₂ network vibrations, thus suggesting an increase in the level of Ta-O-Si bonding. A similar trend can be observed for the xerogels after heating to 750 °C, with the exception of the (Ta₂O₅)_0.16(SiO₂)₀.₈₂ sample which contains far fewer hydroxyl groups. The variation in intensity at 960 cm⁻¹ with Ta

![Fig. 7 Ta L₃-edge EXAFS data for (Ta₂O₅)₀.₄₈(SiO₂)₀.₅₂ after heat treatment to 750 °C: k² weighted EXAFS (top) and Fourier transform (bottom). Experimental data, solid line, and theoretical fit, dotted line.](image1)

![Fig. 8 Ta L₃-edge EXAFS data for (Ta₂O₅)₀.₃₅(SiO₂)₀.₆₅ after heat treatment to 750 °C: k² weighted EXAFS (top) and Fourier transform (bottom). Experimental data, solid line, and theoretical fit, dotted line.](image2)

Table 4 Ta L₃-edge EXAFS derived structural parameters for (Ta₂O₅)ₓ(SiO₂)₁₋ₓ xerogels

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<th>Composition</th>
<th>Heat treatment/°C</th>
<th>Shell</th>
<th>N</th>
<th>R/A</th>
<th>d/A²</th>
<th>R_φ (%)</th>
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<td>Ta-O</td>
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<td>0.011(1)</td>
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<td>(Ta₂O₅)₀.₀₆(SiO₂)₀.₉₄</td>
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</tbody>
</table>

concentration is not so pronounced for the heat treated xerogels. This suggests that for the higher concentrations not all of the Ta atoms are incorporated into the silica matrix. A possible reason for this is partial phase separation of the two components.

This hypothesis is supported by a comparison of the FTIR spectrum from the α-Ta2O5 xerogel with those from the (Ta2O5)x/2(SiO2)1-x samples. The band at 650 cm⁻¹ in the spectrum from the α-Ta2O5 sample is clearly present in the spectra from the (Ta2O5)x/2(SiO2)1-x xerogels both before and after heating to 750 ºC. This suggests that the (Ta2O5)x/2(SiO2)1-x xerogel contains a form of tantalum oxide which is very similar to α-Ta2O5, and hence provides evidence of phase separation in the highest Ta content xerogel. The IR spectra from the (Ta2O5)x/10(SiO2)9/10 samples also exhibit a weak feature at 650 cm⁻¹ which suggests that there is some degree of phase separation present in these samples, although not to the extent of the (Ta2O5)x/2(SiO2)1-x samples.

Opt MAS NMR shows two well defined, distinct oxygen coordination in the pure α-Ta2O5 gel. With heat treatment there are at first some small changes in the relative proportions of these two environments. Also the chemical shifts of these two environments change which is probably related to subtle changes in the exact nature (e.g., the average bond length and/or angles) of these sites as the sample is heated. From the known trends of chemical shift it is possible that the peaks at ~270 and ~430 ppm correspond to local coordinations by n and n-1 tantalum atoms respectively. One of these environments is probably the same as occurs in the crystalline material where there are 34 oxygen environments of which approximately 60% can be designated OTa4 and the remaining OTa3.10 On this basis we can assign the resonances at ~270 and ~430 ppm as those due to OTa4 and OTa3 configurations, respectively. The large difference in the intensities of the OTa3 and OTa2 peaks in the spectrum from the crystalline 750 ºC sample where the two environments are known to be present in approximately 60/40 proportions, respectively, is probably due to a difference in ¹⁷O quadrupolar coupling constants in the two environments. This observed behaviour in the pure oxide has some interesting similarities and differences to ¹⁷O MAS NMR of the amorphous pure oxides TiO2 gels in a mixture of OTa4 and OTa3, with heat treatment reducing the amount of OTa4 until at the point of crystallisation it is eliminated altogether. In TiO2 crystallisation is accompanied by a large reduction in the linewidth which is associated with a large decrease in the disorder on the OTa3 sites.11 Similar narrowing was also observed in the two sites of the monoclinic HfO2 that forms on crystallisation.12 The reduction in disorder on crystallisation in TiO2 and HfO2 reduces the chemical shift distribution that dominates the linewidth in the amorphous gel. Here the linewidth does not decrease much as the Ti2O3 xerogel crystallises, thus reflecting the high degree of atomic disorder in the crystalline state of Ta2O5.

The ¹⁷O MAS NMR results from enriched (Ta2O5)x/20(SiO2)1/20 and (Ta2O5)x/24(SiO2)1/24 samples shown in Figs. 4 and 5, respectively, also yield information on the degree of atomic mixing of the two oxides. The spectra from the (Ta2O5)x/20(SiO2)1/20 samples show clear resonances due to Si–O–Si (~0 ppm) and Ta–O–Si (~160 ppm) bonding and an absence of resonances due to Ta–O–Ta bonding. These spectra are typical of those from an atomically mixed-oxide xerogel13,14 and do not show any appreciable change occurring between the unheated sample and that preheated to 750 ºC. The spectra from the (Ta2O5)x/24(SiO2)1/24 samples exhibit two additional resonances at ~285 and ~445 ppm which have been ascribed to OTa3 and OTa2 environments, respectively. The observation and identification of these resonances is consistent with the conclusion drawn from the FTIR study that partial phase separation occurs within the (Ta2O5)x/25(SiO2)1/25 samples. The only significant change in these spectra with heat treatment of the sample up to 750 ºC is that the OTa3 peak becomes more intense relative to the OTa2 peak. This change is similar to that observed in the spectra from the pure Ta2O3 xerogels and suggests that the environment of the oxygen in the phase separated part of the (Ta2O5)x/20(SiO2)1/20 sample becomes more like that in crystalline Ta2O3 with heat treatment. Despite phase separation, a strong resonance is still observed at 160 ppm for the (Ta2O5)x/24(SiO2)1/24 samples indicating that a significant proportion of the tantalum is atomically mixed with the silica. Similar behaviour has been observed in (ZrO2)1-x(SiO2)x xerogels.15

The Ta L₂,₃-edge EXAFS results presented here provide both detailed structural information of the Ta environment in an atomically mixed (Ta2O5)x/20(SiO2)1/20 xerogel and yield qualitative information on the process of phase separation within this mixed oxide system. Concentrating first on the structural parameters obtained for the Ta–O correlation given in Table 4, it can be seen that on average for all samples, the Ta occupies an environment surrounded by five O atoms at a distance of 1.92 Å. This distance is in agreement with that of 1.92 Å predicted for five coordinate TaVI in an oxide using bond-valence parameters.26 For all the samples, the nearest-neighbour environment of Ta does not appear to change with temperature. However, a structural change is observed across the composition range with the Debye–Waller factor, A, rising from ~0.007 Å² for the (Ta2O5)x/20(SiO2)1/20 samples to ~0.013 Å² for the (Ta2O5)x/24(SiO2)1/24 samples. This change reflects an increase in disorder in the Ta–O nearest-neighbour shell as the concentration of Ta is increased and is most probably due to partial phase separation which is accompanied by the presence of Ta with a coordination number of greater than 5. This observation is consistent with the results from the ¹⁷O MAS NMR study which show the presence of OTa3 configurations in the (Ta2O5)x/25(SiO2)1/25 samples. The presence of OTa4 must be accompanied by an increase in the average Ta–O coordination number, since from bond valence considerations, the presence of only five coordinate Ta would require only OTa3 configurations. The structural parameters in Table 4 obtained from fitting the second nearest-neighbour contributions to the EXAFS signal are less informative due to multiple scattering contribution at the second-shell distances, which have not been accounted for in this analysis, the parameters obtained should be treated as qualitative. The results for all samples suggest that the Ta second-nearest-neighbour is Si at a distance of ~3.2 Å. This distance gives a Ta–O–Si angle of 137° which is entirely reasonable for a material which is predominantly an SiO2 network with Ta dissolved in it, the Si–O–Si bond angle in silicon being 146°.14,27 Again, as the samples were with the highest loading of Ta, no significant structural evolution is observed in the second nearest-neighbour environment of Ta with increasing temperature of the heat treatment. The case of the (Ta2O5)x/23(SiO2)1/25 samples is different. Examining the EXAFS data for the two extremes of composition, shown in Figs. 7 and 8, one can see there is a clear difference between the envelopes of the two sets of oscillations. Looking at the Fourier transforms one can observe a difference above 3 Å. In Fig. 8, for a (Ta2O5)x/25(SiO2)1/25 xerogel heated to 750 ºC, the match between the experimental and calculated Ta–Si peaks in the Fourier transform is poor, with the calculated peak being too wide. It is likely that the fitting procedure is trying to include contributions from other shells, possibly Ta–Ta, within this peak. Attempts to fit a Ta–Ta shell to the EXAFS data were unsuccessful, resulting in non-physical parameters. A similar situation occurred in the fitting of the data from the α-Ta2O5 samples where no meaningful structural parameters could be derived for the Ta–Ta correlations. It remains most probable that the poor modelling of the Ta–Si shell for the (Ta2O5)x/25(SiO2)1/25 samples is due to a contribution from:

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Ta–Ta, resulting from a degree of phase separation, which cannot be modelled without the use of multiple scattering calculations (a highly non-trivial exercise for an amorphous, two-phase, material).

The Ta L\textsubscript{\text{III}}-edge EXAFS derived structural parameters for the \(a\)-Ta\(_2\)O\(_3\) shown in Table 3 offer an insight into the environment of Ta in this relatively poorly characterised material. For the unheated sample and that heated to 250°C, the average coordination of Ta is close 5 with respect to oxygen with a Ta–O distance of 1.93 Å, whereas for the sample heated to 750°C, the Ta–O coordination number and distance are closer to 6 and 2 Å, respectively. The latter result is expected since the XRD results show that crystallisation of the \(a\)-Ta\(_2\)O\(_3\) occurs by 750°C. The \(a\)-Ta\(_2\)O\(_3\) samples both exhibit relatively large Debye-Waller factors of >0.015 Å\(^2\) for the Ta–O shells which probably result from the presence of some Ta with a coordination number of greater than 5. This explanation is supported by the observation of O\(_8\)Ta environments with \(^{17}\)O MAS NMR.

The observation that both the \((Ta_2O_3)\(_3\)(SiO_2)\(_{1-x}\) and \(a\)-Ta\(_2\)O\(_3\) xerogels contain predominantly five-coordinated Ta is particularly interesting. This type of coordination of Ta with respect to oxygen is rare: Ba\(_3\)Si\(_3\)Ta\(_2\)O\(_{12}\) was the first reported solid state compound with five-coordinate tantalum.\(^{28}\) However, five-coordinated Ta is perhaps more likely in amorphous network type structures, particularly in the case of the mixed tantalum–silica xerogels where this type of coordination allows both the Ta–O and Si–O bonds to have a bond valence of one. In this way, inclusion of a small amount of tantalum within the silica structure can occur whilst maintaining a network structure.

Conclusions

The first extensive structural characterisation of \((Ta_2O_3)\(_3\)(SiO_2)\(_{1-x}\) and \(a\)-Ta\(_2\)O\(_3\) xerogels is presented here. EXAFS has allowed identification of five coordinate Ta in both these materials. A Ta environment which is extremely rare in the solid state. The FTIR and \(^{17}\)O MAS NMR studies on the mixed oxides show that partial phase separation of the two component oxides starts to occur for \(x\geq0.11\). Despite partial phase separation, the \((Ta_2O_3)_{0.75}(SiO_2)_{0.25}\) still contains extensive Ta–O–Si bonding; as shown by the \(^{17}\)O MAS NMR results. \(^{29}\) Si MAS NMR data have shown that the presence of Ta within the silica network reduces the connectivity of that network.

The result that partial phase separation can be detected in the \((Ta_2O_3)\(_3\)(SiO_2)\(_{1-x}\) xerogels using a widely available laboratory technique such as FTIR is particularly important because of the detrimental effect that segregation of the two component oxides has on the usefulness of these materials, can be routinely monitored.

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